

## Thermal Stability and Kinetics of Sodium Alginate and Lignosulphonic Acid Blends

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**Abstract:** Biodegradable polymer blends are prepared by mixing two natural biodegradable polymers Sodium alginate (SA) and Lignosulphonic acid (LS) by solution casting method for controlled drug delivery application. Successful application of blend for controlled drug delivery requires its stability under physiological condition. Thermal analysis (TGA) reveal thermal transitions, degradation processes and thermal stability of blends and gives an important insight about its stability. In this paper TGA data has been used to evaluate degradation of SA/LS blends with temperature. The energy of activation for the thermal degradation of polymers and blends has been calculated by 'Horowitz and Metzger' method. These studies reveal that blends are stable as compared to their polymers.

**Keywords:** Sodium alginate, Lignosulphonic acid, Polymer blends, Thermal stability

### 1. INTRODUCTION

Biopolymers are used extensively for wound enclosures, body implants, tissue engineering materials, drug delivery matrices or vehicles, dental applications, diagnostics and therapeutic imaging [1]. These polymers are biocompatible, biodegradable and can be obtained universally. Biopolymers, in their original form may lack in certain features necessary for a particular application. In order to improve properties, and make them suitable for application, they are usually blended with another polymer. Blending of existing polymers can often be achieved rapidly than creating entirely new polymers. The study of polymer's properties, their interaction and resulting features when blended is an important area of study.

Sodium Alginate (SA), a biopolymer obtained from brown sea weed, is extensively used in medicine for wound dressing, cell and enzyme immobilization, controlled drug delivery etc [2]. Its polymeric chain consists of  $\beta$ -D-mannuronic acid (M) and its C-5 epimer  $\alpha$ -L-guluronic acid (G) monomers, which are linked by 1-4 glycosidic linkages. The structure and properties of alginates depends on the G:M ratio, but in general they are biodegradable, non-toxic, biocompatible and have the ability to form hydrogels under mild conditions [3]. Alginate hydrogels are employed for drug

release applications. These hydrogels show poor mechanical stability under physiological conditions [4-5] which is improved by adding suitable polymer/additive as filler. Some of the blends of SA prepared by solvent casting method are listed in Table 1. Aloe Vera gel, known for therapeutic properties was blended with SA by solvent casting method by Ruben Periera et al [6]. Synthesized blend films showed better thermal and mechanical properties than SA and the Film showed stability in simulated body fluid condition. Aloe Vera quantity was found to control the stability and degradation of blend film.

A better water vapor permeability, swelling capacity, thermal stability and tensile strength was achieved by blending SA with Silk fibroin [7]. SEM and XRD of the blend revealed microscopic phase separation and amorphous nature respectively. Chitosan is extensively used in food and medicine for its diverse properties. Blend films with chitosan also showed better thermal and mechanical stability as compared to SA [8]. In order to combine the properties of both chitosan and sodium alginate a blend with medical clay, montmorillonite (MMT) was prepared by solvent casting method also showed improvement in properties and diffusion and swelling controlled drug release [9]. A stable, pH dependent controlled drug release material was achieved by blending Guar gum, sodium alginate and

**Table 1.** List of Sodium alginate (SA) blends

| Sl.No | SA Blends                             | Scope of Study                               | Ref No. |
|-------|---------------------------------------|----------------------------------------------|---------|
| 1.    | SA and Aloe Vera                      | Wound healing and drug delivery applications | 6       |
| 2.    | SA and silk fibroin (SF)              | Wound healing                                | 7       |
| 3.    | SA and Chitosan                       | Improvement in properties                    | 8       |
| 4.    | SA, Chitosan and Montmorillonite      | Drug delivery                                | 9       |
| 5.    | SA and Guar gum and with Cloisite 30B | Drug delivery                                | 10      |
| 6.    | SA and Chitosan                       | Drug delivery                                | 11      |
| 7.    | SA and Collagen                       | Food Packaging                               | 12      |
| 8.    | SA and Poly ethylene glycol(PEG)      | Improvement in properties                    | 13      |
| 9.    | SA and Poly ethylene oxide(PEO)       | Improvement in properties                    | 14      |

Cloisite 30B [10]. In another study viscous and macroscopically homogeneous SA chitosan hydrogel structures were developed by Dominika Kulig et al [11]. The blend showed improvement in thermal, chemical and mechanical properties and stability as compared to SA. TGA analysis showed that thermal decomposition of blend improved by increasing chitosan dosage. In case of SA/collagen, blend films with improved viscosity, thermal stability and mechanical property as compared to pristine polymers was achieved. [12]. Formation of thermally stable and immiscible blend of SA with Poly ethylene glycol was reported by Mruthyunjaya swamy and Ramraj [13]. Poly ethylene oxide (PEO)/ Sodium Alginate (SA) blends showed effect of SA quantity on their thermal degradation temperatures. The thermal behavior of blend was attributed to the hydrogen bond formed between the ether oxygen of PEO and the hydroxyl groups of SA [14].

Inspired by the volumes of literature available on sodium alginate blends and the need for better controlled drug release vehicles, we blended SA with sodium lignosulphonate (LS), a biopolymer from plant source [15]. LS is often used as super plasticizer. The SA/LS blend prepared in 80/20 ratio was found to be most suitable for controlled release of drug Ciprofloxacin under intestinal condition [16]. We have reported in our previous publications

from the XRD and FTIR that the SA/LS blends have shown improved crystallinity while the polymers retained their structure in the blend [17]. Thermal studies are important tool to investigate a material for its suitability as drug delivery vehicle. Polymeric materials with glass transition temperature more than body temperature (37 °C) are stable under aqueous condition and thus are suitable for drug release applications [18]. A material used for controlled release is expected to be stable for longer period of time in aqueous medium in order to retain the drug in its matrix and release the drug systematically in a controlled manner [19]. The knowledge of thermal stability and kinetic parameters of thermal degradation process of polymers and their blends are important in the processing of polymeric material for drug release applications. In the present work, we have investigated the kinetics of thermal degradation of the prepared blends of SA with LS to reveal the effect of blend ratio on thermal stability.

## 2. EXPERIMENTAL

### 2.1. Chemicals

Both SA and LS of analytical grade are purchased from Sigma-Aldrich, India Limited

and used without any further purification. Solutions of polymers have been prepared in distilled water.

## 2. 2. Preparation of SA / LS Blends

SA of desired weight has been soaked in double distilled water for 24 hours prior to mixing. In order to prepare blends of varying composition aqueous solution of thus prepared SA was mixed with aqueous solution of LS in the ratio 100/0, 80/20 and 60/40 by weight and the mixture was continuously stirred for 30 mins using magnetic stirrer. Thin films of different compositions were casted by pouring the resultant mixture on glass substrates, dried at room temperature for 72 h, and then dried at 60 °C in a vacuum oven till constant weight is achieved [20].

## 2. 3. Characterization Techniques

The SA, LS and their blends samples were subjected for TGA in order to investigate the initial and final degradation temperature. The TGA was performed on a TA instrument "Universal V4.7 TA". The sample tests were conducted in an N<sub>2</sub> purge over a temperature range of 20–600 °C at a scan rate of 10 °C/min[21].

## 3. RESULTS AND DISCUSSIONS

In a heating process, a material loses moisture attached to its surface, water entrapped within and decomposes due to breaking of its bonds. It is well established that blending may greatly influence the thermal stability of the component polymers [22-24]. During the process of degradation, considerable interactions may occur between components in the blend and/or their degradation products. The type of interaction depends on the degree of miscibility of the components, nature of the compositions and their ratio in the blend. In general, the first heating stage of polymers due to loss of water show a small (5–15%) mass loss while the second decomposition step show significant mass loss (39–75%). The delay in this process is indicative of stability of a material. Thermal degradation process (TGA) of a material is studied by observing mass loss of a sample as a function of time or temperature in a varied and controlled atmosphere.

The weight loss (TGA) and derivative of weight loss (DTG) curves of pure SA, pure LS and SA/LS blends in ratio 80/20 and 60/40 are presented in Fig. 1 and Fig. 2 respectively. SA in

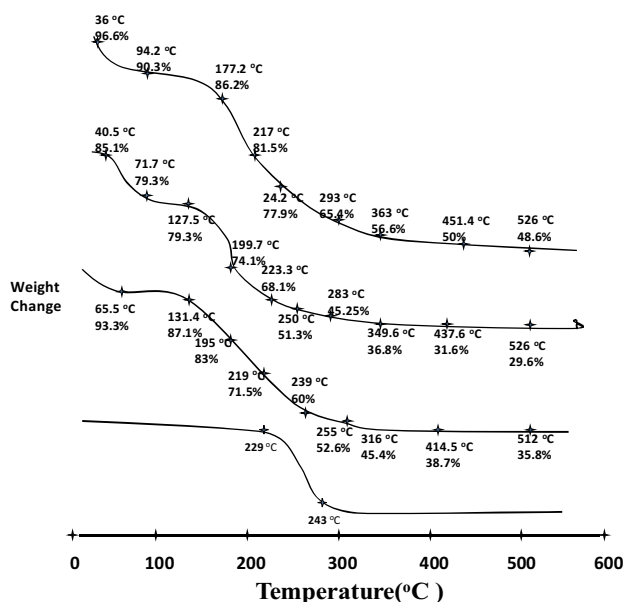


Fig. 1. TGA curves of a) Pure SA b) SA/LS (80/20) c) SA/LS (60/40) and d) Pure LS

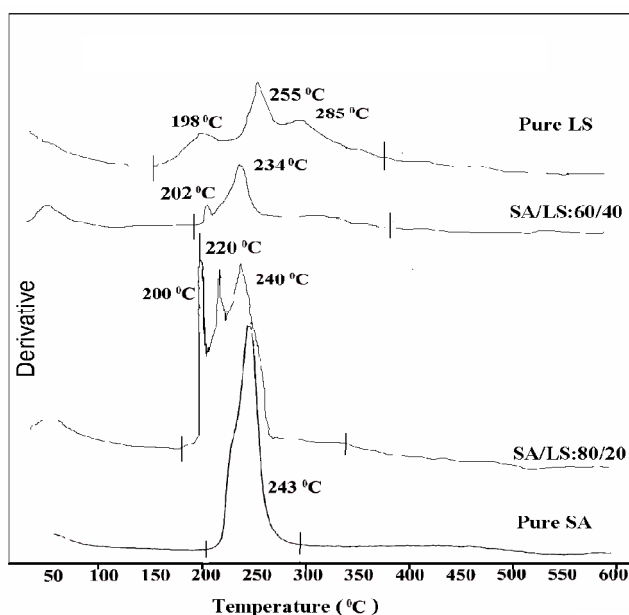


Fig. 2. Derivative of TGA curves of a) Pure SA b) SA/LS (80/20) c) SA/LS (60/40) and d) Pure LS

N<sub>2</sub> atmosphere degrades in two stages - The initial step is due to loss of water attached and the final due to decarboxylation reaction of alginate, fracture of glycosidic bonds, and dehydration and decarboxylation of alginate [21, 25]. Pure LS on the other hand degrades in multiple steps. Its initial mass loss due to water evaporation is observed at relatively lower temperature. Further degradation process is slow and observed over a range of temperature involving chain degradation [26]. TGA and DTG of blends on the other hand show that addition of LS improves thermal stability of SA. The SA/LS (80/20) blend shows degradation over a range of temperature. A lower concentration of LS is expected to reduce

existing intermolecular forces in SA polymeric chains thereby improving the flexibility and chain mobility. The TGA of SA/LS (60/40) blend shows rapid mass loss in the second stage as compared to SA/LS (80/20). This is an indicative of reduced stability due to phase separation. Stability of blends is also reflected by studying the residual mass at 300 °C, 400 °C and 500 °C temperatures [Table 2]. SA/LS (80/20) blends shows a delay in mass loss as compared to the pure SA and the SA/LS (60/40) blend. It can be concluded that the increased mass ratio of LS (40%) exceeds the compatibility limit of SA thus making the blend unstable. This blend has also shown a fast degradation under aqueous condition [16].

Table 2. Degradation temperatures of SA, LS and their blends

| Polymer      | IDT    | Wt. loss (%)<br>300 °C | Wt. loss (%)<br>400 °C | Residue (%)<br>500 °C |
|--------------|--------|------------------------|------------------------|-----------------------|
| Pure SA      | 200 °C | 61.2                   | 68.72                  | 28.38                 |
| SA/LS(80/20) | 195 °C | 51.78                  | 60.68                  | 36.3                  |
| SA/LS(60/40) | 199 °C | 58.39                  | 68                     | 29.8                  |
| Pure LS      | 177 °C | 37.9                   | 47.7                   | 49.39                 |

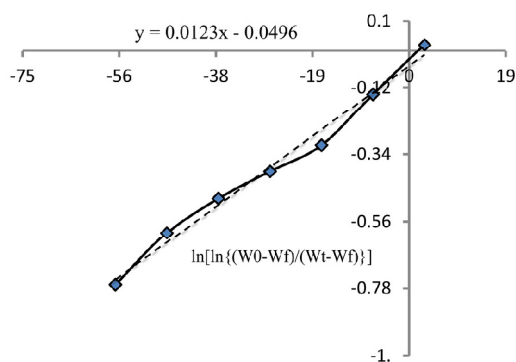


Fig. 3. Calculation of Energy of Activation for pure SA

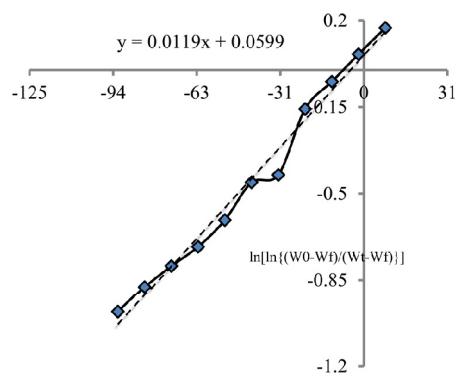


Fig. 4. Calculation of Energy of Activation for blend SA/LS (80/20).

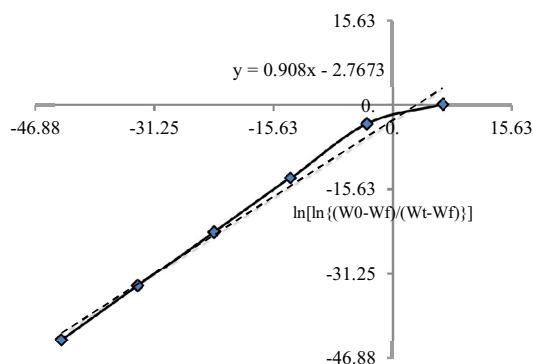


Fig. 5. Calculation of Energy of Activation for blend SA/LS (60/40).

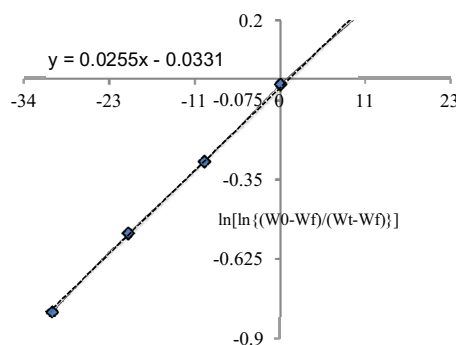


Fig. 6. Calculation of Energy of Activation for pure LS

The thermal data also indicate that the blend formation improves the stability but does not cause change in polymer degradation mechanism. It is clear that the degradation is delayed as the blend has better alternative ways like increase in interfacial area, diffusion of chains among the phases to dissipate the accumulated energy during the heating process. These alternative pathways for the dissipation of the energy increase the thermal stability of polymers [27]. In order to understand the degradation process, energy of activation ( $E_a$ ) for the thermal degradation of pure SA, LS and their blends in the regions 200 °C to 250 °C has been estimated using 'Horowitz and Metzger' method [28]. According to this method a plot of double logarithm of the reciprocal of the reactant versus temperature gives the energy of activation of pyrolysis. For first order kinetics the equation is:

$$\ln[\ln \{(W_0 - W_f) / (W_t - W_f)\}] = E_a \Theta / RT_s^2 \quad (1)$$

where  $W_0$  = initial weight of the sample

$W_f$  = final weight of the sample

$W_t$  = the weight of the sample at temperature 't'.

$R$  = gas constant

$\Theta = T - T_s$ ,  $T_s$  is the reference temperature when  $(W_t - W_f) / (W_0 - W_f) = 1/e$

A plot of double logarithm of the reciprocal of the weight fraction of the reactant versus temperature whose slope ( $E_a \times 10^3$ ) is related to energy of activation.

Fig. 3 to Fig. 6 show the plots obtained from kinetic analysis of TGA data of pure SA, pure LS and their blends SA/LS (80/20) and SA/LS (60/40). The energy of activation is calculated from these plots and results are presented in

**Table 3.** Energy of Activation of pure SA, LS and their blends (SA/LS)

| Samples      | Energy of activation (kJ/mol) |
|--------------|-------------------------------|
| Pure SA      | 6.754                         |
| SA/LS(80/20) | 8.435                         |
| SA/LS(60/40) | 9.261                         |
| Pure LS      | 12.21                         |

Table 3. Energy of activation for pure SA is 6.754 kJ/mol and on blending with LS in the ratio SA/LS (80/20) and SA/LS (60/40) the energy of activation increased to 8.435 and 9.2614 kJ/mol. The increase in  $E_a$  confirms that addition of LS improves the thermal stability of SA.

#### 4. CONCLUSION

TGA and derivative TGA data show that the stability of SA has been enhanced by addition of LS. Study also indicate that the blends are immiscible. Immiscibility leading to two or multiple phases during blending is required in many applications. In case of immiscible blends each phase contributes its own characteristics to the blend material which is a required attribute in some applications. The TGA curves indicate that, SA/LS blends degrade in multiple steps as shown by a multiple peaks. The kinetic studies of SA/LS blends using 'Horowitz and Metzger' method shows better thermal stability than their polymers. The composition 80/20 exhibits better thermal stability as well as the energy of activation thus calculated. It is an established fact that the stability of blends influence the release of drug from polymer matrix and hence the SA/LS blend (80/20) can show better drug release ability and thus have potential applications as controlled drug delivery system.

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